

## LM Reder

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web balancing equations and simple stoichiometry key balance the following equations  $1\text{ N}_2 + 3\text{ F}_2 \rightarrow 2\text{ NF}_3$   $2\text{ C}_6\text{H}_{10} + 17\text{ O}_2 \rightarrow 12\text{ CO}_2 + 10\text{ H}_2\text{O}$   $3\text{ HBr} + 1\text{ KHCO}_3 \rightarrow 1\text{ H}_2\text{O} + 1\text{ KBr} + 1\text{ CO}_2$   $3\text{ GaBr}_3 + 3\text{ Na}_2\text{SO}_3 \rightarrow 1\text{ Ga}_2\text{SO}_3 + 3\text{ SO}_2 + 6\text{ NaBr}$   $3\text{ SnO}_2 + 2\text{ NF}_3 \rightarrow 3\text{ SnF}_2 + 1\text{ N}_2\text{O}_3$  using the following equation  $2\text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$  ideal stoichiometry practice khan academy - Jan 29 2022

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web may 28 2020 write the balanced equation for the reaction and calculate how much  $\text{SiO}_2$  is required to produce 3.00 kg of  $\text{SiC}$  answer  $\text{Ce} + \text{SiO}_2 \rightarrow \text{CeSiO}_3$

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coefficients

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web gabrielle m 9 years ago i m working on chemical reactions double and single replacement on flvs now my first question for this video is how do you have 4 aluminum atoms when it says  $2\text{Al}$  subscript 2 do you just add the 2 and the subscript 2 same thing goes with the  $\text{O}_3$  i am so lost

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web google classroom balance the following chemical equation  $\text{Mg} + \text{OH}^- \rightarrow \text{Mg}(\text{OH})_2$   $\text{HCl} \rightarrow \text{HCl}$   $\text{MgCl}_2 + \text{MgCl}_2 \rightarrow \text{H}_2 + \text{O}_2$  note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem 7 4 1 x x y y  $\theta$   $\pi$  8 5 2 0 9 6

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web a common type of stoichiometric relationship is the mole ratio which relates the amounts in moles of any two substances in a chemical reaction we can write a mole ratio for a pair of substances by looking at the coefficients in front of each species in

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